

Simple Task Card Review

- For typical review, you can simply distribute the Task Cards to your class for individual, group or whole class review
- The cue cards can be projected on the board for whole class review
- The cards work well as cue cards, test review, etc.
- A student worksheet as well as teacher answer key are provided

Task Card Review Game

Instructions

- Students get into groups of two
- Task cards are divided into sets of five where #1-5 is one set, #6-10 is the second, etc.
- Place task cards around the room randomly (i.e. # 1 shouldn't necessarily be near #2)
 - o Note – the game works better if you print multiple copies of each Task Card to place around the room so the stations don't get too crowded.
- Each group is given a task card set to work on first (i.e. Group 1 starts on task card set 1, etc.)
- Since there are five questions per task card set, the members decide who is responsible for each question
- Students disperse, find their questions and answer them on a separate piece of paper
- On an overhead projector, have the Task Card Review Game Word Chart displayed.
 - o Note – printing and posting copies around the room will decrease congestion around the projector and make it easier to read.
- Once the group has completed their task card set, they will have a series of words which when put into the correct order based on task card number (#1, #2, etc.), will make-up a portion of a quote
- The group then brings their series of words to the teacher who checks for correctness
- If the group is correct, they are assigned a new task card set with five new questions and the game continues.

Once the group has completed each task card set and put together the quote, they are the winner.

It may sound complicated but it really isn't. Give it a try and you'll find it's very straightforward.

Student Answer Sheet

1 -	11 -	21 -
2 -	12 -	22 -
3 -	13 -	23 -
4 -	14 -	24 -
5 -	15 -	25 -
6 -	16 -	
7 -	17 -	
8 -	18 -	
9 -	19 -	

10 -

20 -

10 -	20 -	
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Teacher Answer Key

When playing the Task Card Review Game, the quote that should be completed once all the Task Card Sets are complete is – “The most exciting phrase to hear in science, the one that heralds the most discoveries, is not "Eureka!" (I found it!) but "That's funny..." - Isaac Asimov

1 - $\text{Fe}_2(\text{CO}_3)_3$, tin (III) fluoride	11 - P_2O_5 , carbon tetrabromide	21 - Na_3N , barium oxide
2 – gains electrons.	12 - $\text{K}_2\text{SO}_4 + \text{H}_2\text{O}$	22 - lead (IV) oxide + calcium \rightarrow calcium oxide + lead
3 - $4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$	13 - $\text{AgNO}_{3(aq)} + \text{NaCl}_{(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{3(aq)}$	23 - Neutralization
4 – non-metal	14 – $\text{Al}(\text{OH})_3(l)$	24 - ionic compound
5 – protons and neutrons	15 - Orange Juice \rightarrow Clean Rainwater \rightarrow Salt Water \rightarrow Baking Soda \rightarrow Ammonia	25 - $2\text{NaBr} + \text{Ca}(\text{OH})_2 \rightarrow 2\text{NaOH} + \text{CaBr}_2$
6 - SiO_2 , ammonium cyanide	16 - Li_3As , magnesium chloride	
7 - $2\text{AlBr}_3 + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3 + 3\text{Br}_2$	17 - $4\text{Al} + 3\text{PbBr}_4 \rightarrow 4\text{AlBr}_3 + 3\text{Pb}$	
8 – 5 grams.	18 - CO_2 and H_2O	
9 – $\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + \text{H}_2$	19 – weak base	
10 – $[\text{H}^+$ ion]	20 - 4, 1, 5, 2, 3	

Task Card Review Game Word Chart

Answer	Word	Answer	Word	Answer	Word
Li ₃ As, magnesium chloride	is	$4\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_{3(s)}$	is	Fe(CO ₃), tin (III) fluoride	have
Fe ₂ (CO ₃) ₃ , tin (III) fluoride	“The	polar compound.	infinite	ionic compound.	Isaac
CO ₂ and H ₂ O	“Eureka!”	gains electrons.	most	Single Replacement	and
Si ₂ O, ammonium cyanide	are	2.5 grams.	the	[H ⁺ ion]	one
P ₂ O ₅ , monocarbon tribromide	Adams	CO and C	the	$\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + \text{H}$	true
loses electrons.	universe	Fe ₂ (CO ₃), tin (I) fluoride	and	$3\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + \text{H}_2$	not
lead (IV) oxide + calcium → calcium oxide + lead	‘That’s	SiO ₂ , nitrogen hydrogen ammonium cyanide	the	$\text{Pb}_2 + \text{H}_3\text{PO}_4 \rightarrow \text{Pb}_2(\text{PO}_4)_2 + \text{H}_2$	sure
ions.	things	3LiAs, magnesium chloride	the	H ₂ O _(l)	Two
SO _{2(g)}	fundamental	$4\text{Al} + 3\text{PbBr}_4 \rightarrow 4\text{AlBr}_3 + 3\text{Pb}$	not	$2\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{(s)}$	Carl
5 grams.	science,	$\text{Al} + 3\text{Pb}_4\text{Br} \rightarrow \text{AlBr}_3 + 3\text{Pb}$	we	$\text{NaBr} + \text{Ca}(\text{OH})_2 \rightarrow \text{NaOH} + \text{CaBr}_2$	can
$4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$	exciting	$2\text{AlBr}_3 + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3 + 3\text{Br}_2$	in	SiO, ammonium cyanide	of
Clean Rainwater → Salt Water → Ammonia → Lye → Baking Soda → Orange Juice	the	Orange Juice → Clean Rainwater → Salt Water → Baking Soda → Ammonia → Lye	discoveries,	Clean Rainwater → Lye → Ammonium → Baking Soda → Orange Juice → Salt Water	human
2 grams.	not	Li ₃ As, magnesium dichloride	science	non-metal.	phrase
alkali metal.	mysterious	$2\text{Al}_2 + 2\text{Pb}_2\text{Br} \rightarrow 2\text{AlBr}_3 + 2\text{Pb}$	experience	PO, monocarbon tetrabromide	Sagan
$\text{Pb} + \text{H}_3\text{PO}_4 \rightarrow \text{Pb}_3(\text{PO}_4)_2 + \text{H}_2$	the	$\text{NaBr}_2 + \text{Ca}(\text{OH})_2 \rightarrow \text{Na}_2\text{OH} + \text{CaBr}_2$	art	SiO ₂ , ammonium cyanide	hear
Al(OH) _{3(aq)}	most	metalloid.	Rutherford	$2\text{AlBr}_2 + \text{Cl}_2 \rightarrow 2\text{AlCl} + \text{Br}$	true
gains protons.	I’m	$\text{AgNO}_3_{(aq)} + \text{NaCl}_{(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_3_{(aq)}$	the	$2\text{NaBr} + \text{Ca}(\text{OH})_2 \rightarrow 2\text{NaOH} + \text{CaBr}_2$	Asimov
Na ₃ N, barium oxide	but	$\text{AlBr}_3 + 3\text{Cl}_2 \rightarrow \text{AlCl}_3 + \text{Br}_2$	beautiful	weak base.	(I found
[OH ⁻ ion]	emotion	number of protons.	Albert	protons and neutrons.	to
P ₂ O ₅ , carbon tetrabromide	that	gains electrons.	most	KSO ₄ + H ₂ O	cradle
K ₂ SO ₄ + H ₂ O	heralds	neutral.	Douglas	1, 4, 5, 2, 3	Bohr
$\text{AgNO}_2_{(aq)} + \text{NaCl}_3_{(aq)} \rightarrow 3\text{AgCl}_{(s)} + 3\text{NaNO}_{(aq)}$	which	electrons and neutrons.	and	4, 1, 5, 2, 3	it!)
Na ₃ N, barium dioxide	stupidity	$\text{AgNO}_{(aq)} + \text{NaCl}_{(aq)} \rightarrow \text{AgCl}_{(s)} + \text{NaNO}_{(aq)}$	stands	Neutralization	funny...’

Task Card Set #1:

Question	Answer	Word
1 - Write the formula/name for the following: •iron (III) carbonate •SnF ₃	Fe ₂ (CO ₃) ₃ , tin (III) fluoride	“The
2 - An atom becomes a negatively charged ion when it	gains electrons.	most
3 - For the following, write the chemical equation with states and balance. Solid aluminum metal reacts with oxygen from the air to form a protective solid coating called aluminum oxide.	$4\text{Al}_{(s)} + 3\text{O}_{2(g)} \rightarrow 2\text{Al}_2\text{O}_{3(s)}$	exciting
4 - A bright yellow solid substance is known to be an element. It is brittle and will shatter if hit with a hammer. It is a poor conductor of both heat and electricity. This element is best called a(n)	non-metal.	phrase
5 - The mass number of an atom is the sum of its	protons and neutrons.	to

Task Card Set #2

Question	Answer	Word
<p>6 - Write the formula/name for the following:</p> <ul style="list-style-type: none"> • silicon dioxide • NH₄CN 	SiO ₂ , ammonium cyanide	hear
<p>7 - Balance the following chemical equation:</p> <p>__ AlBr₃ + __ Cl₂ → __ AlCl₃ + __ Br₂</p>	2AlBr ₃ + 3Cl ₂ → 2AlCl ₃ + 3Br ₂	in
<p>8 - 23 grams of magnesium reacts with 12 grams of hydrogen chloride to produce 30 grams of magnesium chloride and a gas. What is the mass of the gas?</p>	5 grams.	science,
<p>9 - Predict the product(s) for the following reaction (assume Pb²⁺):</p> <p>Pb + H₃PO₄ →</p>	Pb + H ₃ PO ₄ → Pb ₃ (PO ₄) ₂ + H ₂	the
<p>10 – What determines the acidity of a substance</p>	[H ⁺ ion]	one

Task Card Set #3

Question	Answer	Word
<p>11 - Write the formula/name for the following:</p> <ul style="list-style-type: none"> • diphosphorus pentoxide • CBr₄ 	<p>P₂O₅, carbon tetrabromide</p>	<p>that</p>
<p>12 - The products of a neutralization reaction between sulphuric acid and potassium hydroxide would be:</p>	<p>K₂SO₄ + H₂O</p>	<p>heralds</p>
<p>13 - For the following, write the chemical equation with states and balance:</p> <p>Aqueous silver nitrate reacts with sodium chloride dissolved in water to yield solid silver chloride and aqueous sodium nitrate.</p>	<p>AgNO₃ (aq) + NaCl (aq) → AgCl (s) + NaNO₃ (aq)</p>	<p>the</p>
<p>14 - In the following equation, the "X" represents:</p> <p>Al₂(SO₄)₃(aq) + Ca(OH)₂(aq) → X + CaSO₄(aq)</p>	<p>Al(OH)₃ (l)</p>	<p>most</p>
<p>15 - Take the following list and arrange it from most strongly acidic to most strongly basic.</p> <ul style="list-style-type: none"> • Lye - pH 14 • Salt Water - pH 7 • Orange Juice - pH 4 • Ammonia - pH 12 • Baking Soda - pH 8 • Clean Rainwater - pH 6 	<p>Orange Juice → Clean Rainwater → Salt Water → Baking Soda → Ammonia → Lye</p>	<p>discoveries,</p>

Task Card Set #4

Question	Answer	Word
16 - Write the formula/name for the following: • lithium arsenide • MgCl ₂	Li ₃ As, magnesium chloride	is
17 - Write out the chemical reaction for the following sentence: Aluminum and lead (IV) bromide reacts to form aluminum bromide and lead.	$4\text{Al} + 3\text{PbBr}_4 \rightarrow 4\text{AlBr}_3 + 3\text{Pb}$	not
18 - If a hydrocarbon fully combusts, what are the two products?	CO ₂ and H ₂ O	"Eureka!"
19 - A substance has a pH of 9, the substance is a	weak base.	(I found
20 - Put the following periodic table groups in order from left to right 1-Alkali Earth Metals, 2-Halogens, 3-Noble Gases, 4-Alkali Metals, 5-Transition Metals	4, 1, 5, 2, 3	it!)

Task Card Set #5

Question	Answer	Word
21 - Write the formula or name for the following: • sodium nitride • BaO	Na ₃ N, barium oxide	but
22 - What is the proper word equation for the following reaction: PbO ₂ + 2Ca → 2CaO + Pb	lead (IV) oxide + calcium → calcium oxide + lead	'That's
23 - What type of reaction is HNO ₃ + Ca(OH) ₂ → Ca(NO ₃) ₂ + H ₂ O	Neutralization	funny..."
24 - NaNO ₃ is a(n)	ionic compound.	Isaac
25 - Predict the product(s) for the following reaction and balance: NaBr + Ca(OH) ₂ →	2NaBr + Ca(OH) ₂ → 2NaOH + CaBr ₂	Asimov



1

Write the formula/name for the following:

- iron (III) carbonate
- SnF_3



2

An atom becomes a negatively charged ion when it



3

For the following, write the chemical equation with states and balance.

Solid aluminum metal reacts with oxygen from the air to form a protective solid coating called aluminum oxide.



4

A bright yellow solid substance is known to be an element. It is brittle and will shatter if hit with a hammer. It is a poor conductor of both heat and electricity. This element is best called a(n)

5

The mass number of an atom is the sum of its

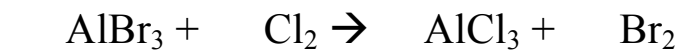
6

Write the formula/name for the following:

- silicon dioxide
- NH_4CN

7

Balance the following chemical equation:



8

23 grams of magnesium reacts with 12 grams of hydrogen chloride to produce 30 grams of magnesium chloride and a gas. What is the mass of the gas?

9

Predict the product(s) for the following reaction (assume Pb^{2+}):



10

What determines the acidity of a substance?

11

Write the formula/name for the following:

- diphosphorus pentoxide
- CBr_4

12

The products of a neutralization reaction between sulphuric acid and potassium hydroxide would be:

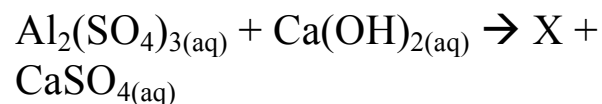
13

For the following, write the chemical equation with states and balance:

Aqueous silver nitrate reacts with sodium chloride dissolved in water to yield solid silver chloride and aqueous sodium nitrate

14

In the following equation, the "X" represents:



15

Take the following list and arrange it from most strongly acidic to most strongly basic.

- Lye - pH 14
- Salt Water - pH 7
- Orange Juice - pH 4
- Ammonia - pH 12
- Baking Soda - pH 8
- Clean Rainwater - pH 6

16

Write the formula/name for the following:

- lithium arsenide
- MgCl_2

17

Write out the chemical reaction for the following sentence:

Aluminum and lead (IV) bromide reacts to form aluminum bromide and lead.

18

If a hydrocarbon fully combusts, what are the two products?

19

A substance has a pH of 9, the substance is a(n)

20

Put the following periodic table groups in order from left to right

1-Alkali Earth Metals, 2-Halogens, 3-Noble Gases, 4-Alkali Metals, 5-Transition Metals

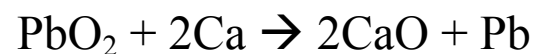
21

Write the formula or name for the following:

- sodium nitride
- BaO

22

What is the proper word equation for the following reaction:



23

What type of reaction is
 $\text{HNO}_3 + \text{Ca}(\text{OH})_2 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$

24

NaNO_3 is a(n)



25

Predict the product(s) for the following reaction and balance:

